4. (b) :

	CO_2	+ C (graphite)	$\rightleftharpoons$	2CO
Initial pressure (atm)	0.5	0		0
At equilibrium (atm)	$(0.5 - x)$	–		$2x$

Total pressure at equilibrium = 0.8 atm

Thus, $0.5 - x + 2x = 0.8$

$x = 0.3$ atm

$$K_p = \frac{(P_{CO})^2}{P_{CO_2}} = \frac{(2 \times 0.3)^2}{0.2} = 1.8 \text{ atm}$$