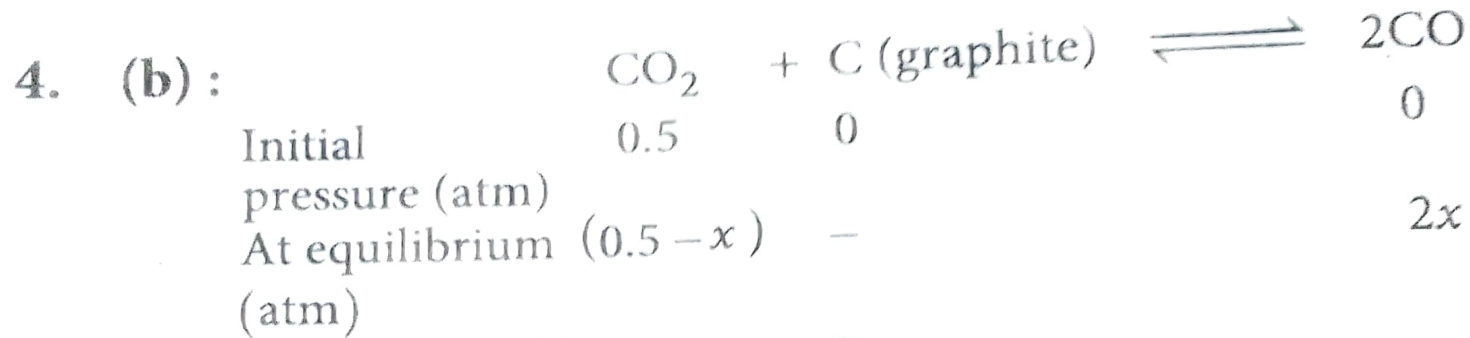




Total pressure at equilibrium = 0.8 atm

Thus, $0.5 - x + 2x = 0.8$

$$x = 0.3 \text{ atm}$$

$$K_p = \frac{(P_{\text{CO}})^2}{P_{\text{CO}_2}} = \frac{(2 \times 0.3)^2}{0.2} = 1.8 \text{ atm}$$