(b) : Electrolysis of concentrated sulphuric acid produces peroxydisulphuric acid ($H_2S_2O_8$) along with O_2 and H_2.

At cathode : $H^+_{(aq)} + e^- \longrightarrow \frac{1}{2}H_2$

At anode : $2SO^{2-}_{4(aq)} \longrightarrow S_2O^{2-}_{8(aq)} + 2e^-$

$S_2O_8^{2-}$ combines with H^+ ions to produce $H_2S_2O_8$.