

(d) : For isoelectronic species, ionic radii increase with decrease in magnitude of nuclear charge.

(A  $\rightarrow$  IV)

Ionisation energy increases along a period. However, ionisation energy of N  $>$  O due to stable half-filled electronic configuration of N. (B  $\rightarrow$  I)

Metallic character increases down the group and decreases along a period. (C  $\rightarrow$  II)

Electronegativity increases across a period. (D  $\rightarrow$  III)